Chemistry H.W

1] Calculate the amount of 1 M NaOH aqueous solution needed to make 100 mL of 0.5 M NaOH aqueous solution.

2] The element boron consists of two isotopes, $^{10}_{5}B$ and $^{11}_{5}B$. Their masses, based on the carbon scale, are 10.01 and 11.01, respectively. The abundance of $^{10}_{5}B$ is 20.0%.
What is the atomic abundance of and the abundance of $^{11}_{5}B$?

3] Avogadro's Law Problem

A 6.0 L sample at 25°C and 2.00 atm of pressure contains 0.5 mole of a gas. If an additional 0.25 mole of gas at the same pressure and temperature are added, what is the final total volume of the gas?

4] Calculate the mass in grams of a single carbon (C) atom.

1]To prove charle's law temperature is taken in

A. high temperature
B. farenheit scale
C. Kelvin scale
D. centigrade scale

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Coldest temperature which is ever achieved is

A. 0 kelvin
B. 0 degree
C. 0 farenheit
D. none

Slope of a straight line is increased in graph of Charle's law by increasing

A. mass of gas
B. pressure of gas
C. temperature of gas
D. moles

Lowest temperature for a substance to be in gaseous state is

A. atmosphere
B. 273.16 degree negative
C. 208 K
D. 0 degree centigrade

When volume becomes zero temperature of gases is

A. 0 kelvin
B. 0 degree centigrade
C. 0 farenheit
D. none

A macromolecule found in blood is

A. Ferritin
B. albumin
C. Haemoglobin
D. Keratin

. Atom is composed just of electrons and protons this concept prevailed till

A. 1980
B. 1932
C. 1970
D. 1960

. Important part of atom which is related to many physical and chemical properties is its

A. size
B. mass
C. nucleus
D. protons

. A molecule is a smallest particle which can exist

A. independently
B. in combination
C. in space
D. in gas form

. Units by which hydrogen atom is lighter than haemoglobin is

A. 70000
B. 58000
C. 78000
D. 68000